

Name: _____

Pd: _____ Date: _____

Chemistry 12
Acid Base Part II
Titration Curve Worksheet

1. Write the **net ionic equation** for the following titrations:

a. NH_3 and HBr



b. H_2SO_4 and RbOH

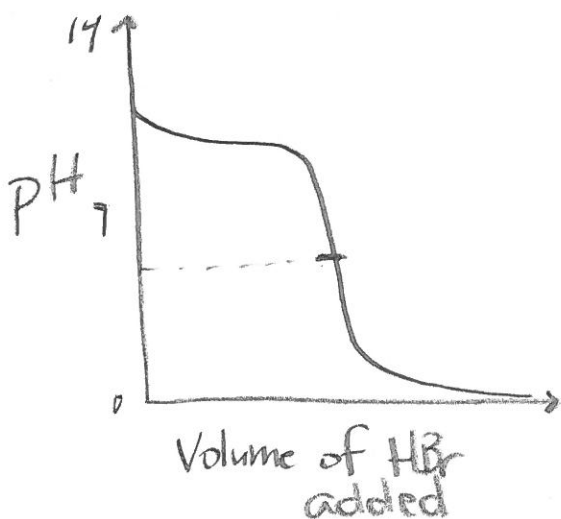


c. CH_3COOH and KOH

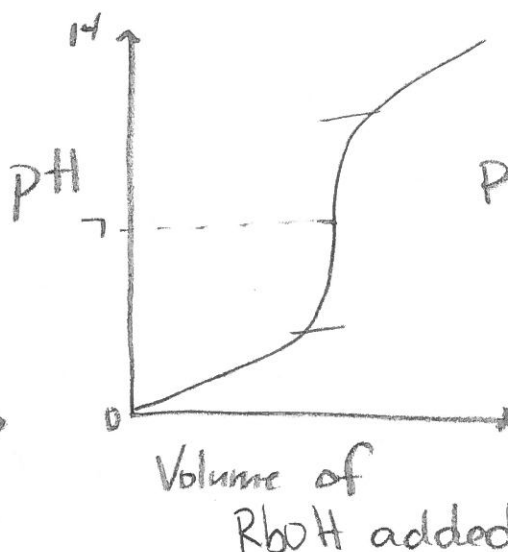


2. For each of the **above reactions**, draw the **titration curve** that would represent each. Be sure to have a title, and labeled x and y axis

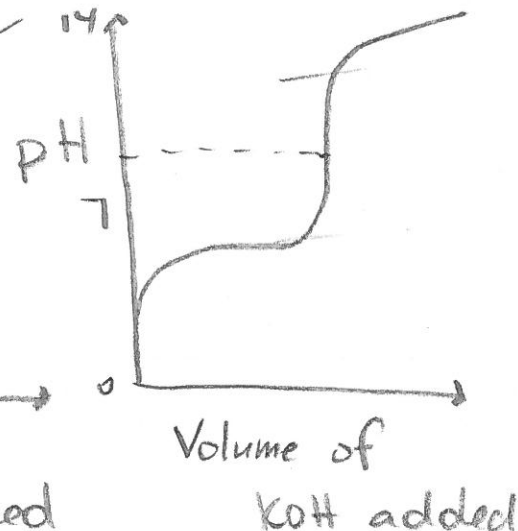
a.



b.



c.



3. Given the indicators: Methyl red, Phenol red and Thymol blue, choose the indicator that would be used in the above titrations.

a.

methyl red
pKa = 5.4

b.

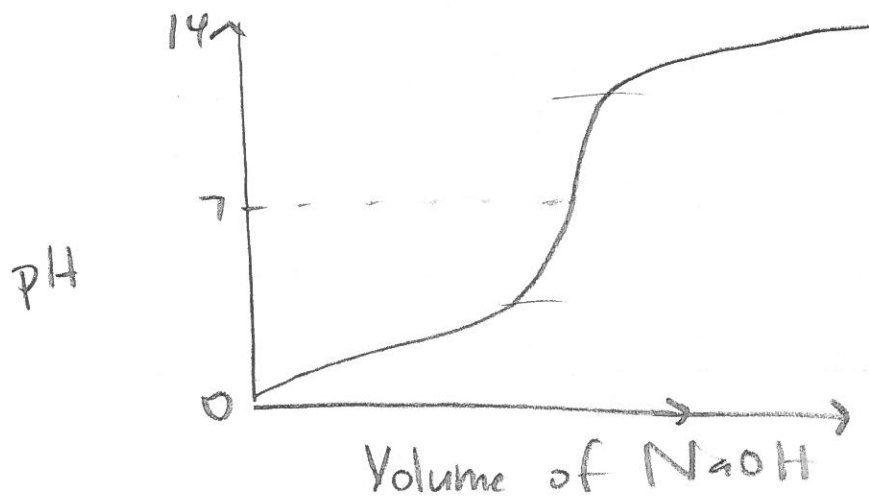
phenol blue
pKa = 7.3

c.

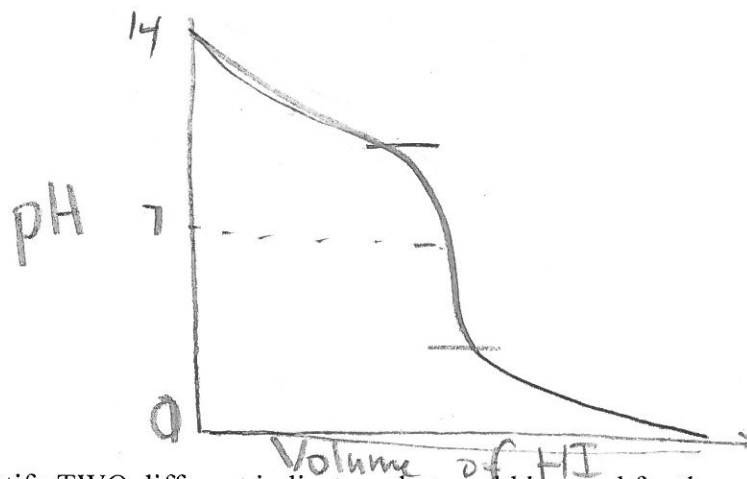
Thymol blue
pKa = 8.8

8.0
9.6

4. For the titration between HI and NaOH, draw the two versions of the titration curve:
a. pH vs Volume of base added



- b. pH vs Volume of acid added



- c. Identify TWO different indicators that could be used for these titrations:

neutral red $pK_a = 7.4$

Bromthymol blue $pK_a = 6.9$