Name:_____
Blk:_____Date:_____

Chemistry 12
Solubililty Lesson #4
Writing Formula, Complete and Net Ionic Equations

In this unit, every reaction that produces a precipitate wil be a _____! Recall from Chemistry 11 that this is a situation in which the _____!!!

Example 1: Write the reaction occuring when 0.2 M solutions of $Ca(NO_3)_2$ and Na_3PO_4 are mixed.

Step 1. Identify the REACTANTS

Step 2. Identify the IONS present in solution

Step 3. Identify the PRODUCTS, by switching partners

Step 4. Look on table to identify product STATES

Step 5. BALANCE the equation

Once the balanced equation is detemind there are THREE POSSIBLE ways to WRITE THE BALANCED CHEMICAL REACTION

1. FORMULA EQUATION:

2. COMPLETE IONIC EQUATION:

3. NET IONIC EQUATION:

Example 2.
Write a formula equation, complete ionic equation and a net ionic equation for the following reactions in which:

a. 0.2 M $Pb(NO_3)_2$ and 0.2 M $FeCl_3$ are mixed

b. 0.2 M Silver nitrate and 0.2 M Sodium phosphate

SEAT WORK/ HOMEWORK: Exercise 25 (odd letters only)
PLO's: G7, H1, H2 AND H3