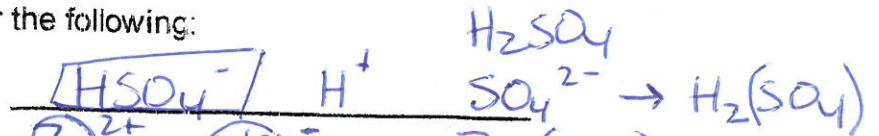


Chemical Formulas POLYATOMICS

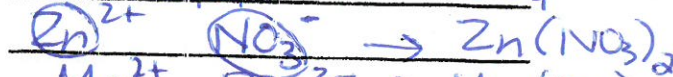
Name _____
Date _____ Blk _____

1. Write the chemical formula for the following:

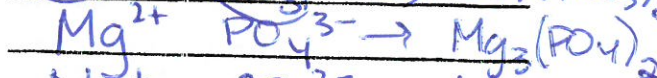
a. hydrogen sulphate



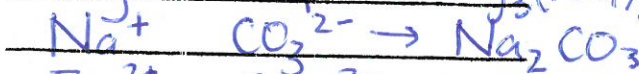
b. zinc nitrate



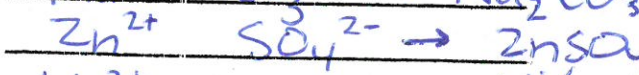
c. magnesium phosphate



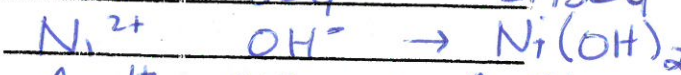
d. sodium bicarbonate



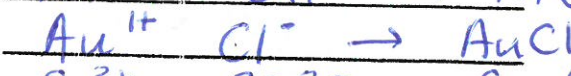
e. zinc sulphate



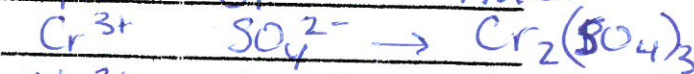
f. nickel (II) hydroxide



g. gold (I) chloride



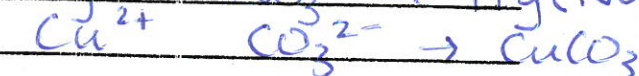
h. chromium (III) sulphate



i. mercury (II) nitrate



j. copper (II) bicarbonate



2. Write the chemical name for the following:

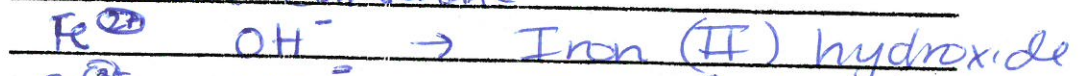
a. CuSO_4



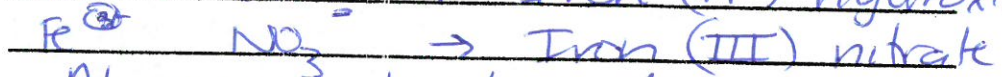
b. Na_2CO_3



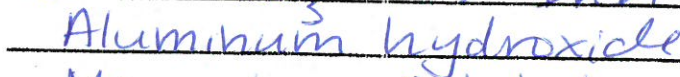
c. $\text{Fe}(\text{OH})_2$



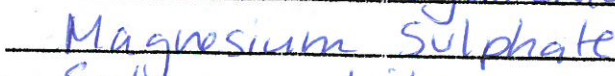
d. $\text{Fe}(\text{NO}_3)_3$



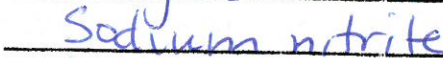
e. $\text{Al}(\text{OH})_3$



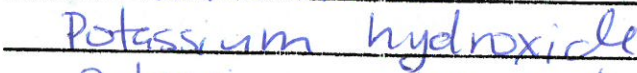
f. MgSO_4



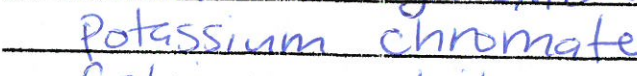
g. NaNO_2



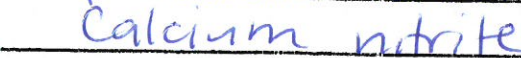
h. KOH



i. K_2CrO_4



j. $\text{Ca}(\text{NO}_2)_2$

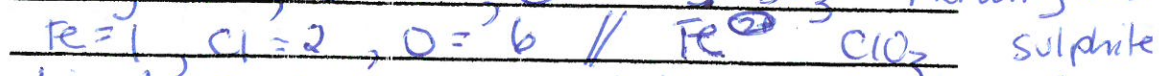


3. Indicate the number of atoms of each element.

a. HgSO_3



b. $\text{Fe}(\text{ClO}_3)_2$



c. LiOH



d. $\text{Pb}_2(\text{CO}_3)_4$



e. $\text{Mn}(\text{OH})_2$



Iron (II) chlorate

Lead (IV) carbonate

Naming Compounds

Name Key
Block _____ Date _____

A. 2 Element Compounds:

- a) ScF_3
- b) AgI
- c) MgCl_2
- d) ZnO
- e) Al_2O_3

Scandium fluoride

Silver iodide

Magnesium chloride

Zinc oxide

Aluminum oxide

B. Metallic Elements with 2 Combining Capacities

- a) CuO
- b) PbS
- c) PbI_2
- d) Au_2O_3
- e) SnO_2

Copper (II) oxide

Lead (II) sulphide

Lead (II) iodide

Gold (III) oxide

Tin (IV) oxide

C. Compounds with Polyatomic Ions

- a) KClO_3
- b) $\text{Mg}(\text{OH})_2$
- c) $\text{Cr}(\text{NO}_3)_3$
- d) $\text{Mn}(\text{SO}_4)$
- e) $\text{Co}_3(\text{PO}_3)_2$

Potassium chlorate

Magnesium hydroxide

Chromium (III) nitrate

Manganese (II) sulphate

Cobalt (II) phosphite

D. Figure it out for yourself!!!!

- a) FePO_4
- b) Fe_2O_3
- c) AuCl_3
- d) CoCl_3
- e) $(\text{NH}_4)_2\text{CO}_3$
- f) $\text{Al}(\text{OH})_3$
- g) SnO
- h) MgCO_3
- i) NH_4OH
- j) HgCr_2O_7
- k) FeSO_4
- l) AgCl
- m) PbO_2
- n) $\text{Ba}(\text{HCO}_3)_2$
- o) Cu_2S

Iron (II) phosphate

Iron (III) oxide

Gold (III) chloride

Cobalt (III) chloride

Ammonium carbonate

Aluminum hydroxide

Tin (II) oxide

Magnesium carbonate

Ammonium hydroxide

Mercury (II) dichromate

Iron (II) sulphate

Silver chloride

Lead (IV) oxide

Barium hydrogen carbonate

Copper (I) sulphide

 CO_3