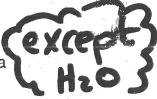
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## Chemistry 11 Notes on Naming Acids

For the purposes of Chemistry 11, any compound that begins with a **HYDROGEN** atom will be considered an **ACID**.



## **RULES FOR NAMING ACIDS:**

- #1. When Hydrogen is paired with a MONATOMIC NON-METAL
  - a. Hydrogen becomes hydro
  - b. Change the non-metal ending to IC
  - c. Add the word ACID

Ex. HCl → <u>hydrochloric acid</u> Ex HBr → <u>hydrobromic acid</u>

#2. When Hydrogen is paired with a **POLYATOMIC NON-METAL** that ends in "ATE"

- a. eliminate Hydrogen from the name
- b. change the ATE ending to IC
- c. add the word ACID

sulphak - sulphuric

Ex.  $H_2SO_4 \rightarrow \underline{\text{sulpt 0 ic acid}}$ Ex.  $H_3PO_4 \rightarrow \underline{\text{phosphoric acid}}$ 

#3. When Hydrogen is paired with a **POLYATOMIC NON-METAL** that ends in "ITE"

- a. eliminate Hydrogen from the name
- b. change the ITE ending to OUS
- c. add the word ACID

Ex. HNO

→ Nitrous acid

Ex. HSO<sub>3</sub>

→ Sulphurous acid

note PO3 = phosph

## **DIATOMIC MOLECULES, THE RULE OF "GEN" and "INE"**

Some elements in there natural state occur as Diatomic molecules, that is they have two atoms in their molecular structure.

These include: N<sub>2</sub>, O<sub>2</sub>, F<sub>2</sub>, Cl<sub>2</sub>, Br<sub>2</sub>, I<sub>2</sub> and H<sub>2</sub>.

not astitine

A little trick to remember them by is that they form the **No. 7** on the periodic table plus **Hydrogen**. (excluding astatine)

sulphur is an octamelecule... Sa.