Name:_____
Blk:_____Date:_____

Chemistry 11
Notes on Naming Acids

except H_2O

For the purposes of Chemistry 11, any compound that begins with a **HYDROGEN** atom will be considered an **ACID.**

RULES FOR NAMING ACIDS:

#1. When Hydrogen is paired with a **MONATOMIC NON-METAL**

a. Hydrogen becomes **hydro**
b. Change the non-metal ending to **IC**
c. Add the word **ACID**

Ex. HCl → **hydrochloric acid**
Ex HBr → **hydrobromic acid**

#2. When Hydrogen is paired with a **POLYATOMIC NON-METAL** that ends in "**ATE**"

a. **eliminate** Hydrogen from the name
b. change the **ATE** ending to **IC**
c. add the word **ACID**

sulphate → sulphuric

Ex. H_2SO_4 → **sulphuric acid**
Ex. H_3PO_4 → **phosphoric acid**

#3. When Hydrogen is paired with a **POLYATOMIC NON-METAL** that ends in "**ITE**"

a. **eliminate** Hydrogen from the name
b. change the **ITE** ending to **OUS**
c. add the word **ACID**

note PO_3^{3-} = phosphite

Ex. HNO_2 → **Nitrous acid**
Ex. H_2SO_3 → **Sulphurous acid**

DIATOMIC MOLECULES, THE RULE OF "GEN" and "INE"

Some elements in there natural state occur as Diatomic molecules, that is they have two atoms in their molecular structure.

These include: N_2 , O_2 , F_2 , Cl_2 , Br_2, I_2 and H_2. not astitine

A little trick to remember them by is that they form the **No. 7** on the periodic table plus **Hydrogen**. (excluding astatine)

Sulphur is an octamolecule
$\therefore S_8$.