

Name: _____

Blk: _____ Date: _____

Chemistry 11
Notes on Naming Acids

except
H₂O

For the purposes of Chemistry 11, any compound that begins with a HYDROGEN atom will be considered an ACID.

RULES FOR NAMING ACIDS:

#1. When Hydrogen is paired with a MONATOMIC NON-METAL

- Hydrogen becomes hydro
- Change the non-metal ending to IC
- Add the word ACID

Ex. HCl → hydrochloric acid

Ex HBr → hydrobromic acid

#2. When Hydrogen is paired with a POLYATOMIC NON-METAL that ends in "ATE"

- eliminate Hydrogen from the name
- change the ATE ending to IC
- add the word ACID

sulphate → sulphuric

Ex. H₂SO₄ → sulphuric acid

Ex. H₃PO₄ → phosphoric acid

#3. When Hydrogen is paired with a POLYATOMIC NON-METAL that ends in "ITE"

- eliminate Hydrogen from the name
- change the ITE ending to OUS
- add the word ACID

note PO₃³⁻ = phosphite

Ex. HNO₂ → Nitrous acid

Ex. H₂SO₃ → Sulphurous acid

DIATOMIC MOLECULES, THE RULE OF "GEN" and "INE"

Some elements in their natural state occur as Diatomic molecules, that is they have two atoms in their molecular structure.

These include: N₂, O₂, F₂, Cl₂, Br₂, I₂ and H₂.

not astatine

A little trick to remember them by is that they form the No. 7 on the periodic table plus Hydrogen. (excluding astatine)

sulphur is an octamolecule
∴ S₈.