Mixed Naming Worksheet – Ionic, Covalent and Acids

Directions: Be sure to determine what type of compound your substance is – ionic, binary covalent or an acid.

Naming Compounds

1. NO_2 Nitrogen Dioxide
2. NaBr Sodium Bromide
3. SiO_2 Silicon Dioxide
4. P_2Br_4 Diphosphorus Tetrabromide
5. $FeSO_4$ Iron (II) Sulfate
6. SF_6 Sulfur Hexafluoride
7. HNO_3 Nitric Acid
8. Li_2S Lithium Sulfide
9. HCl Hydrochloric acid
10. $MgBr_2$ Magnesium Bromide
11. N_2S Dinitrogen Monosulfide
12. SeF_2 Selenium Difluoride
13. $AsCl_3$ Arsenic Trichloride
14. HClO Hypochlorous Acid
15. $Be(OH)_2$ Beryllium Hydroxide
16. SO_3 Sulfur trioxide
17. $KMnO_4$ Potassium Permanganate
18. $HClO_4$ Perchloric Acid
19. Cu_2S Copper (I) Sulfide
20. BF_3 Boron TriFluoride
21. H_2SO_4 Sulfuric Acid
22. CaI_2 Calcium Iodide
23. $Pb_3(PO_4)_2$ Lead (II) Phosphate
24. HCH_3COO Acetic Acid
25. $Ca(CH_3COO)_2$ Calcium acetate
26. P_2O_5 Diphosphorus Pentaoxide
27. $BaCl_2$ Barium chloride
28. K_2SO_4 Potassium Sulfate
29. $Pb(ClO_3)_2$ Lead (II) Chlorate
30. H_3PO_4 Phosphoric Acid
31. NF_3 Nitrogen Trifluoride
32. $TiCl_2$ Titanium (II) Chloride
33. $Al_2(HPO_4)_3$ Aluminum Hydrogen Phosphate
34. $HBrO_3$ Bromic Acid
35. HIO_3 Iodic Acid
36. H_2S Hydrosulfuric Acid
37. Si_3Cl_9 Trisilicon Nonachloride
38. Ca_3P_2 Calcium Phosphide
39. $Sr(OH)_2$ Strontium Hydroxide
40. $HClO_2$ Chlorous Acid
41. $HClO_4$ Perchloric Acid
42. H_2SO_3 Sulfurous Acid
43. H_3P Hydrophosphoric Acid
44. NO Nitrogen monoxide
45. HBr Hydrobromic Acid
46. $HClO_2$ chlorous Acid
47. $Mg(MnO_4)_2$ Magnesium Permanganate
48. P_4Cl_8 Tetraphosphorus Octachloride
49. Ca_3N_2 Calcium Nitride
50. SnS_2 Tin (IV) Sulfide

Writing Formulas

51. hydroiodic acid HI
52. hyrdrosulfuric acid H_2S
53. calcium sulfide CaS
54. dinitrogen pentoxide N_2O_5
55. aluminum sulfate $Al_2(SO_4)_3$
56. sulfurous acid H_2SO_3
57. nitric acid HNO_3
58. dihydrogen monoxide H_2O
59. trisulfur monochloride S_3Cl
60. selenium monoxide SeO
61. hydrotelluric acid H_2Te
62. tin(IV) sulfite $Sn(SO_3)_2$
63. carbonic acid H_2CO_3
64. barium acetate $Ba(C_2H_3O_2)_2$
65. zinc(II) periodate omit
66. chloric acid $HClO_3$
67. silicon pentanitride SiN_5
68. bromic acid $HBrO_3$
69. pentaphosphorous heptoxide P_5O_7
70. sodium oxide Na_2O
71. hydrofluoric acid HF
72. calcium bromate $Ca(BrO_3)_2$
73. hyrdrobromic acid HBr
74. silicon dioxide SiO_2
75. nickel(III) sulfide Ni_2S_3
76. manganese(II) phosphate $Mg_3(PO_4)_2$
77. silver(I) iodate KIO_3
78. hydrobromic acid HBr
79. diboron tetrabromide B_2B_4
80. phosphoric acid H_3PO_4
81. potassium carbonate K_2CO_3
82. ammonium oxide $(NH_4)_2O$
83. carbon monoxide CO
84. aluminum sulfite $Al_2(SO_3)_3$
85. zinc(II) nitrate $Zn(NO_3)_2$
86. sodium carbonate Na_2CO_3
87. sodium permanganate Na_2MnO_4
88. diphosphorus pentoxide P_2O_5
89. lead(IV) sulfide PbS_2
90. copper(I) sulfate Cu_2SO_4
91. aluminum sulfate $Al_2(SO_4)_3$
92. ammonium nitride $(NH_4)_3N$
93. magnesium hydroxide $Mg(OH)_2$
94. calcium cyanide $Ca(CN)_2$
95. nitric acid HNO_3
96. sulfuric acid H_2SO_4
97. xenon tetrafluoride XeF_4
98. pentaphosphorus hexafluoride P_5F_6
99. cobalt(II) hypochlorite $Co(ClO)_2$
100. aluminum bicarbonate $Al(HCO_3)_3$