

WORKSHEET #7: NOMENCLATURE OF HYDRATED IONIC COMPOUNDS

Use Latin prefixes to indicate # of water molecules present

1 = mono 2 = di 3 = tri 4 = tetra 5 = penta
6 = hexa 7 = hepta 8 = octa 9 = nona 10 = deca

#	Name of Hydrate	Chemical Formula
Eg.	copper(II)sulfate pentahydrate	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
1		$\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$
2	sodium carbonate decahydrate	
3		$\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$
4	barium chloride dihydrate	
5		$\text{Cd}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$
6		$\text{ZnCl}_2 \cdot 6\text{H}_2\text{O}$
7	zinc sulfate heptahydrate	
8	lithium chloride tetrahydrate	
9		$\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$
10	cobalt(II)chloride hexahydrate	
11		$\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$
12		$\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$
13	barium hydroxide octahydrate	
14	nickel(II)chloride hexahydrate	
15		$\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$
16	iron(III)phosphate tetrahydrate	
17		$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
18	calcium sulfate dihydrate	
19		$\text{SnCl}_4 \cdot 5\text{H}_2\text{O}$
20	barium bromide tetrahydrate	

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#	Name of Hydrate	Chemical Formula
Eg.	copper(II)sulfate pentahydrate	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
1	Magnesium sulphate hepta-hydrate	$\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$
2	sodium carbonate decahydrate	$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$
3	Magnesium chloride hexahydrate	$\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$
4	barium chloride dihydrate	$\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$
5	Cadmium nitrate tetrahydrate	$\text{Cd}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$
6	Zinc chloride hexahydrate	$\text{ZnCl}_2 \cdot 6\text{H}_2\text{O}$
7	zinc sulfate heptahydrate	$\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$
8	lithium chloride tetrahydrate	$\text{LiCl} \cdot 4\text{H}_2\text{O}$
9	Sodium thiosulphate pentahydrate	$\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$
10	cobalt(II)chloride hexahydrate	$\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$
11	Aluminum chloride hexahydrate	$\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$
12	Calcium chloride dihydrate	$\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$
13	barium hydroxide octahydrate	$\text{Ba}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$
14	nickel(II)chloride hexahydrate	$\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$
15	Sodium sulphate decahydrate	$\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$
16	iron(III)phosphate tetrahydrate	$\text{FePO}_4 \cdot 4\text{H}_2\text{O}$
17	Iron (II) sulphate heptahydrate	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
18	calcium sulfate dihydrate	$\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$
19	Tin (IV) chloride pentahydrate	$\text{SnCl}_4 \cdot 5\text{H}_2\text{O}$
20	barium bromide tetrahydrate	$\text{BaBr}_2 \cdot 4\text{H}_2\text{O}$

