

## WORKSHEET #7: NOMENCLATURE OF HYDRATED IONIC COMPOUNDS

Use Latin prefixes to indicate # of water molecules present

1 = mono      2 = di      3 = tri      4 = tetra      5 = penta  
6 = hexa      7 = hepta      8 = octa      9 = nona      10 = deca

#	Name of Hydrate	Chemical Formula
Eg.	copper(II)sulfate pentahydrate	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
1		$\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$
2	sodium carbonate decahydrate	
3		$\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$
4	barium chloride dihydrate	
5		$\text{Cd}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$
6		$\text{ZnCl}_2 \cdot 6\text{H}_2\text{O}$
7	zinc sulfate heptahydrate	
8	lithium chloride tetrahydrate	
9		$\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$
10	cobalt(II)chloride hexahydrate	
11		$\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$
12		$\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$
13	barium hydroxide octahydrate	
14	nickel(II)chloride hexahydrate	
15		$\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$
16	iron(III)phosphate tetrahydrate	
17		$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
18	calcium sulfate dihydrate	
19		$\text{SnCl}_4 \cdot 5\text{H}_2\text{O}$
20	barium bromide tetrahydrate	