Name:_____

Partner:_____

Science 9 Fruit Lab

Read through pages 276 to 277 in *BC SCIENCE 9* before attempting this experiment

Objective:

1. To create a battery using 2 different fruits as the **electrolyte source** and pairs of different metals as the **electrodes**

Pre-lab Questions:

****Define the term **electrolyte**:

****Define the term **electrode**:

****Which **fruit** (of the two you and your partner have selected) do you think will work best as an **electrolyte source**?

*****Which pair of metals (from the list provided) do you think will be the best electrodes?

<u>Procedure:</u> As seen on pages 276-277 in <u>BC SCIENCE 9.</u> You will perform PART 1 using two different fruit before proceeding to part 2.

Data and Observations:

	FRUIT #1 :	Metal 2				
Metal 1		Aluminum	Zinc	Iron	Copper	
	Aluminum		X	X	X	
	Zinc			X	X	
	Iron				Х	
	Copper					

	FRUIT #2		Metal 2				
		Aluminum	Zinc	Iron	Copper		
Metal 1	Aluminum		x	x	x		
	Zinc			X	X		
	Iron				X		
	Copper						

Analysis:

1. In Part 1:

- a. Which fruit produced the best voltage results?
- b. For the fruit selected in a, which combination of metals produced the lowest voltage?
- c. For the fruit selected in a, which combination of metals produced the highest voltage?
- 2. In general, how did the voltage produced by two similar metals compare to the voltages produced when two different metals were used?
- 3. In Part 2, how did the voltage produced by the two metals in **pure water** compare to the results you gathered for when the metals were placed in the fruit? Give a possible explanation for this.

Conclude and Apply:

- 1. a. Now that the experiment is concluded, list the components necessary for the highest voltage battery:
 - b. Were the answers to your pre-lab questions supported or rejected?
- 2. a. How is the fruit battery that you made in this experiment similar to a store-bought battery?

b. How is the fruit battery that you made in this experiment different to a store-bought battery?