Name: $\qquad$
Partner: $\qquad$

Science 9 Fruit Lab
Read through pages 276 to 277 in BC SCIENCE 9 before attempting this experiment

## Objective:

1. To create a battery using 2 different fruits as the electrolyte source and pairs of different metals as the electrodes

## Pre-lab Questions:

****Define the term electrolyte:
****Define the term electrode:
****Which fruit (of the two you and your partner have selected) do you think will work best as an electrolyte source?
****Which pair of metals (from the list provided) do you think will be the best electrodes?

Procedure: As seen on pages 276-277 in BC SCIENCE 9. You will perform PART 1 using two different fruit before proceeding to part 2.

## Data and Observations:

| Metal 1 | FRUIT \#1 | Metal 2 |  |  |  |
| :---: | :---: | :---: | :---: | :---: | :---: |
|  |  | Aluminum | Zinc | Iron | Copper |
|  | Aluminum |  | X | X | X |
|  | Zinc |  |  | X | X |
|  | Iron |  |  |  | X |
|  | Copper |  |  |  |  |


| Metal 1 | FRUIT \#2 | Metal 2 |  |  |  |
| :---: | :---: | :---: | :---: | :---: | :---: |
|  |  | Aluminum | Zinc | Iron | Copper |
|  | Aluminum |  | X | X | X |
|  | Zinc |  |  | X | X |
|  | Iron |  |  |  | X |
|  | Copper |  |  |  |  |

## Analysis:

1. In Part 1:
a. Which fruit produced the best voltage results?
b. For the fruit selected in a, which combination of metals produced the lowest voltage?
c. For the fruit selected in a, which combination of metals produced the highest voltage?
2. In general, how did the voltage produced by two similar metals compare to the voltages produced when two different metals were used?
3. In Part 2, how did the voltage produced by the two metals in pure water compare to the results you gathered for when the metals were placed in the fruit? Give a possible explanation for this.

## Conclude and Apply:

1. a. Now that the experiment is concluded, list the components necessary for the highest voltage battery:
b. Were the answers to your pre-lab questions supported or rejected?
2. a. How is the fruit battery that you made in this experiment similar to a store-bought battery?
b. How is the fruit battery that you made in this experiment different to a store-bought battery?
