Excess Quantities key

26. Mass of CS_2 produced = 22.2 g CS_2

Excess of SO_2 ⇒ 39.5 g - 37.4 g = 2.1 g SO_2

27. Mass of NO produced = 26.8 g NO

Excess of Cu ⇒ 87.0 g - 85.0 g = 2.0 g Cu

28. Mass of P_4 produced = 8.06 g P_4

Excess of $Ca_3(PO_4)_2$ ⇒ 41.5 g - 40.3 g = 1.2 g $Ca_3(PO_4)_2$

Excess of SiO_2 ⇒ 26.5 g - 23.4 g = 3.1 g SiO_2

29. Mass of Br_2 produced = 36.9 g Br_2

Excess of $K_2Cr_2O_7$ ⇒ 25.0 g - 22.7 g = 2.3 g $K_2Cr_2O_7$

Excess of H_2SO_4 ⇒ 60.0 g - 52.9 g = 7.1 g H_2SO_4

30. Volume of CO_2 produced = 24.3 L CO_2 (g)

31. NaOH is in excess

32. Mass of $BaBr_2$ formed = 0.279 g $BaBr_2$