

Name: _____

Blk: _____ Date: _____

Chemistry 12
EQUILIBRIUM Lesson #8 E+F
EQUILIBRIUM CALCULATIONS

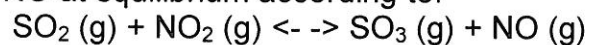
Example Type E:

$K_{eq}=3.5$ for $\text{SO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons \text{SO}_3(\text{g}) + \text{NO}(\text{g})$

If 4.0 mol of $\text{SO}_2(\text{g})$ and 4.0 mol of $\text{NO}_2(\text{g})$ are placed into a 5.0 L bulb and allowed to come to equilibrium, what concentration of all species will exist at equilibrium?

Example Type F:

A 1.0 L reaction vessel contained 1.0 mol of SO_2 , 4.0 mole of NO_2 , 4.0 mol of SO_3 and 4.0 mol of NO at equilibrium according to:



If 3.0 mol of SO_2 is added to the mixture, what will be the new concentrations of NO when equilibrium is re-attained?

SEAT WORK/HOMEWORK: Exercises 55-65 pgs 71-72

PLO F8 (PLEASE NOTE that we have NOT COVERED: PLO's D2, D3 and E3)