Chemistry 12
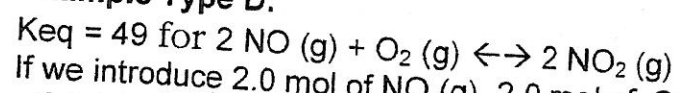EQUILIBRIUM Lesson #8 C+D
EQUILIBRIUM CALCULATIONS

Example Type C:
A certain amount of NO_2 (g) was introduced into a 5.00 L bulb. When equilibrium
was attained according to the equation:

$$2 NO (g) + O_2 (g) \leftrightarrow 2 NO_2 (g)$$
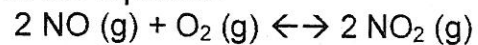
the concentrations of NO (g) was 0.800 M. If the Keq has a value of 24.0 how
many moles of NO_2 were originally put into the bulb?

Example Type D:

Keq = 49 for $2 NO_{(g)} + O_{2 (g)} \leftrightarrow 2 NO_{2 (g)}$

If we introduce 2.0 mol of $NO_{(g)}$, 2.0 mol of $O_{2 (g)}$ and 0.40 mol of $NO_{2 (g)}$ into a 2.0 L bulb, which will the reaction shift in order to reach equilibrium?

SEAT WORK/HOMEWORK: Exercises 47- 54 pgs 70-71
PLO's: F5,F6 and F7