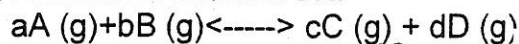


Name: _____

Blk: _____ Date: _____

Chemistry 12
EQUILIBRIUM Lesson #6
THE EQUILIBRIUM EXPRESSION AND K_{eq}

Given the following **EQUILIBRIUM EQUATION**:



where the lower case letters represent _____ and the upper case letters represent _____

We can write out an **EQUILIBRIUM EXPRESSION** (aka K_{eq}) the numerical value of the K_{eq} is called the EQUILIBRIUM CONSTANT

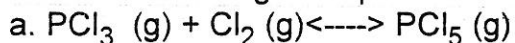
The equilibrium expression is conventionally written as:



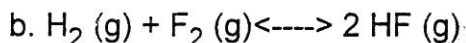
Therefore the above generic equation is written as:

$$K_{eq} =$$

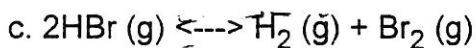
Write the following examples out as **EQUILIBRIUM EXPRESSIONS** :



$$K_{eq} =$$



$$K_{eq} =$$

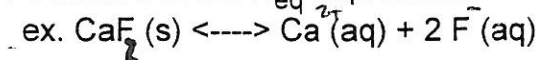


$$K_{eq} =$$

NOTICE THAT THE ABOVE CHEMICALS ARE ALL IN THEIR _____
PHASES! THIS IS IMPORTANT BECAUSE GENERALLY _____
_____ SPECIES ARE INCLUDED IN THE EQUILIBRIUM EXPRESSION

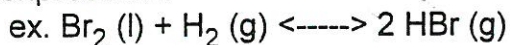
WHY? K_{eq} only includes species whose *concentrations* can CHANGE:

SOLIDS cannot be appreciably compressed (volume can't be decreased) therefore they cannot change their MOLAR CONCENTRATIONS (mol/L). This is why SOLIDS are NOT included in the K_{eq} expression

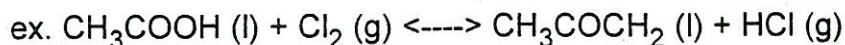


$$K_{eq} =$$

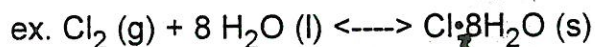
LIQUIDS also cannot be appreciably compressed. HOWEVER, if there is ANOTHER LIQUID present it CAN cause both liquids to become DILUTE and, therefore, change both liquid's concentrations. This is why if the balanced equilibrium equation contains a SINGLE liquid it is NOT included in the K_{eq} expression.



$K_{eq} =$



$K_{eq} =$



$K_{eq} =$

TO SUMMARIZE: Because the concentrations of solids and pure liquids cannot change, adding them as a reactant or product to a system in equilibrium will have NO EFFECT. That is, the equilibrium does NOT SHIFT. THEREFORE; when writing a K_{eq} expression, solids and SINGLE liquids are NEVER included whereas gases, aqueous and multiple liquids are.

SEAT WORK/HOMEWORK: Exercises 31- 35 pg 60
PLO's: F2