Name:_____
Blk:_____Date:_____

Chemistry 12
EQUILIBRIUM Lesson #4
Le Chatelier's Principle

Le Chatelier's Principle STATES:_____
_____
_____

IN OTHER WORDS: **"Whatever We DO, NATURE tries to UNDO"**

We will use the following chemical equation throughout this lesson:
$2\ NO\ (g) + Cl_2\ (g) \leftrightarrows 2\ NOCl\ (g) + 76\ kJ$

1. **Effect of TEMPERATURE CHANGES:**

_____
_____

a. AN INCREASE IN TEMPERATURE for the above equilibrium reaction will result in a _____.

b. A DECREASE IN TEMPERATURE for the above equilibrium reaction will result in a _____.

GRAPHICALLY:

2. **Effect of CONCENTRATION CHANGES**:

_____
_____

a. AN INCREASE IN [Cl_2] for the above equilibrium reaction will result in a _____.

b. A DECREASE IN [Cl_2] for the above equilibrium reaction will result in a _____.

GRAPHICALLY:

3. Effect of PRESSURE CHANGES:

_____

_____

a. AN INCREASE IN PRESSURE for the above equilibrium reaction will result in a _____.

b. A DECREASE IN PRESSURE for the above equilibrium reaction will result in a _____.

GRAPHICALLY:

4. Effect of ADDING A CATALYST:

_____

_____

GRAPHICALLY:

Seatwork/Homework: Exercises 17 – 28 pgs 54-55
PLO's :E1, E2, E4