

Name: _____

Blk: _____ Date: _____

Chemistry 12
EQUILIBRIUM Lesson #2

On the given piece of graph paper answer exercises #6+7 pgs 40-41 HEBDEN.
This will be collected at the end of the class for marks.

CHARACTERISTICS OF EQUILIBRIUM

The system that you are investigating is at equilibrium when:

1. The RATE of _____ is EQUAL to the RATE of _____

2. When the reactants concentration is _____ with time and
the products concentration is _____ with time.

NB: the [REACTANTS] ≠ [PRODUCTS]

3. The forward and reverse _____ DO NOT _____ as time
passes.

4. A closed system that is not in _____ will tend to move toward
a position of _____.

5. The FORWARD RATE _____ the REVERSE RATE



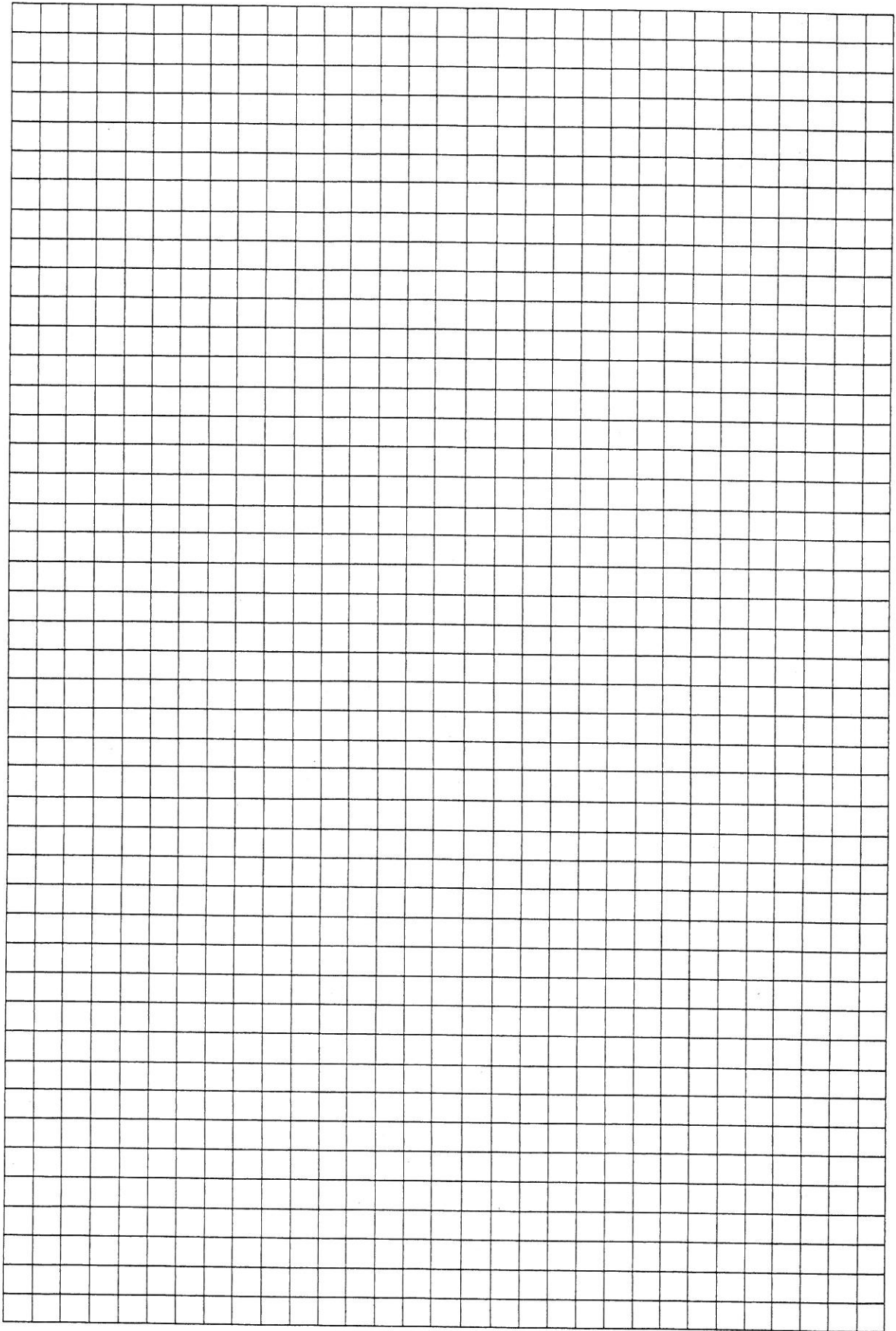
6. When the above example is in equilibrium:

_____ mole of H_2 and _____ mole of Br_2 react for every _____ moles of HBr
that is produced. While _____ moles of HBr reacts for every _____ mole of H_2
and _____ mole of Br_2 that is produced.

SEAT WORK/HOMEWORK: Exercises 8-13 pgs 42-43

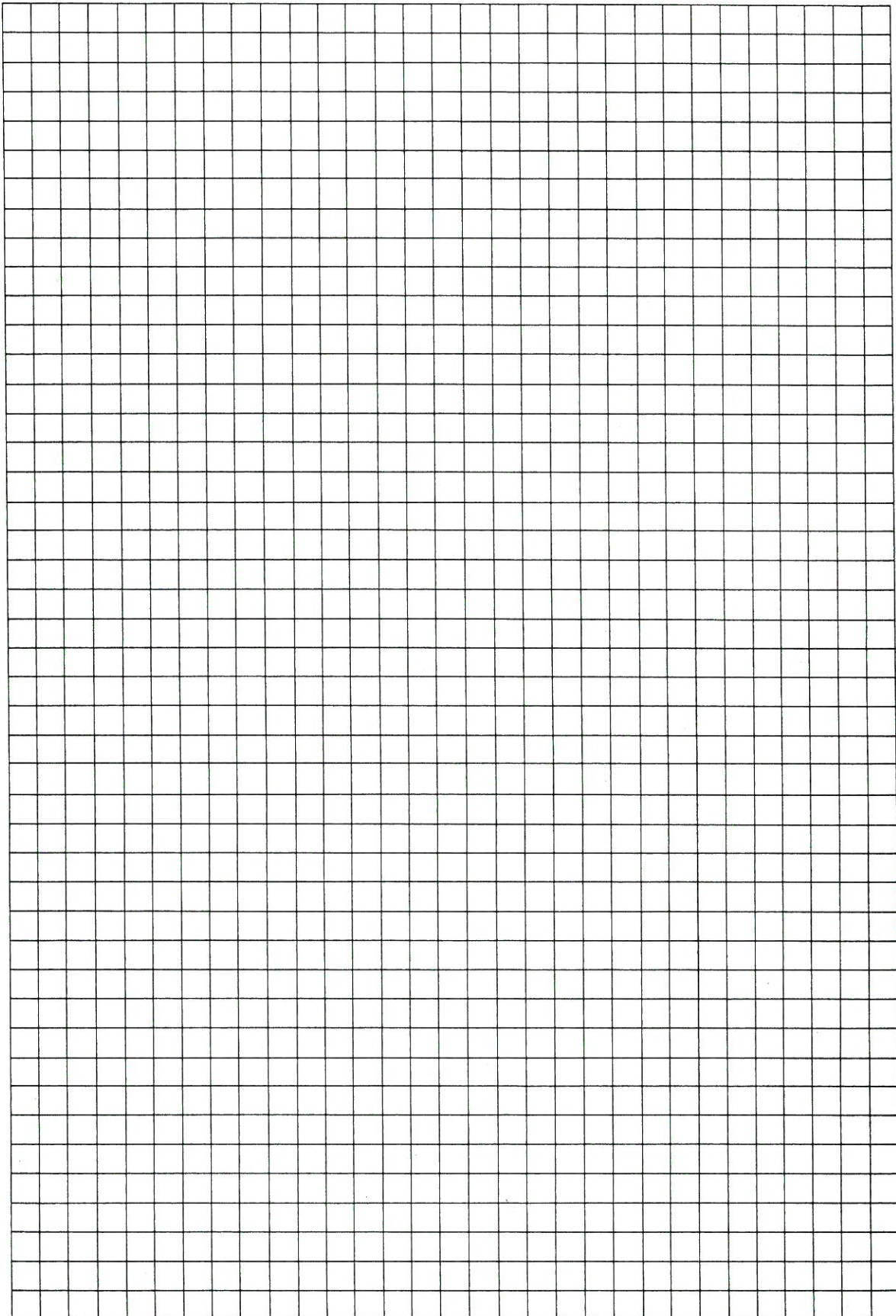
PLO's: D3 and D6 + F1 (graph)

Name _____



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