

Name: _____

Blk: _____ Date: _____

Chemistry 12
Electrochemistry Lesson #5
Balancing Redox Equations using Half-Reactions

IMPT: There are TWO ways to Balance REDOX Equations: using Half-Reactions and using oxidation numbers....you learn both.

This technique balances the redox equation by _____

_____.

Example 1: Balance $\text{Os} + \text{IO}_3^- \rightarrow \text{OsO}_4 + \text{I}_2$ (in acidic solution)

Step 1:

Step 2:

Step 3:

Step 4:

Example 2. Balance $\text{MnO}_4^- + \text{C}_2\text{O}_4^{2-} \rightarrow \text{MnO}_2 + \text{CO}_2$ (in basic solution)

Step 1:

Step 2:

Step 3:

Step 4:

Step 5:

Example 3: Balance $\text{ClO}_2^- \rightarrow \text{ClO}_3^- + \text{Cl}^-$ (in basic solution)

Step 1:

Step 2:

Step 3:

Step 4:

Step 5:

DISPROPORTIONATION: _____

Seatwork/Homework: do exercise 24 EVEN LETTERS ONLY!!!!

PLO's T2 and T3