

Name: _____

Blk: _____ Date: _____

Chemistry 11 Class Starter

A. HYDRATES

Name the following:

- a. $Ba(HCO_3)_2 \cdot 6H_2O$ Barium bicarbonate hexahydrate
- b. $AuCl_3 \cdot 4H_2O$ Gold III chloride tetrahydrate
- c. $Ni_2(CrO_4)_3 \cdot 3H_2O$ Nickel III chromate trihydrate
- d. $LiMnO_4 \cdot 2H_2O$ Lithium permanganate dihydrate

Write the formula for the following:

- a. Iron (III) nitrite heptahydrate $Fe(NO_2)_3 \cdot 7H_2O$
- b. Tin (II) sulphate monohydrate $SnSO_4 \cdot H_2O$
- c. Calcium sulphide octahydrate $CaS \cdot 8H_2O$
- d. Potassium sulphite trihydrate $K_2SO_3 \cdot 3H_2O$

B. ACIDS

Name the following :

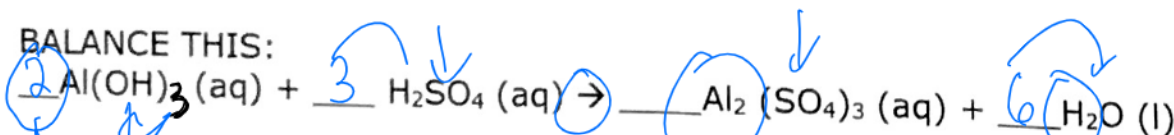
- a. HF hydrofluoric acid
- b. H_2CO_3 carbonic acid
- c. HClO hypochlorous acid
- d. H_2S hydrosulphuric acid

Write the formula for the following :

- a. Hydronitric acid H_3N
- b. nitrous acid HNO_2
- c. Permanganic acid $HMnO_4$
- d. Hydroselenic acid H_2Se

P.A. $\begin{matrix} \text{Metals} \\ \text{Ions} \\ \text{Non-metals} \\ \text{Hydrogen} \\ \text{Oxygen} \end{matrix}$

BALANCE THIS:



H 6

2 = Al = 2

3 = SO₄ = 3