Name:_____
Blk:____ Date:_____

Electrochemistry
Class Starter for Lesson 7

If $KMnO_4$ and $Fe(NO_3)_2$ react in the presents of sulfuric acid (H_2SO_4), $Mn(NO_3)_2$ and $Fe_2(SO_4)_3$ are formed.

Write the **dissociation equation** for $KMnO_4$:

Write the **dissociation equation** for $Fe(NO_3)_2$

Write the **dissociation equation** for $Mn(NO_3)_2$

Write the **dissociation equation** for $Fe_2(SO_4)_3$:

Now balance the following **net ionic REDOX reaction** under **ACIDIC** conditions:

$$\mathbf{Mn}O_4^{1-} + \mathbf{Fe}^{2+} \rightarrow \mathbf{Mn}^{2+} + \mathbf{Fe}^{3+}$$