

Name: _____

Blk: _____ Date: _____

Chemistry 11
Titration Lab

1. Use a funnel and pour the 0.50 M NaOH into the buret
2. Use suction bulb and pipet to measure out 10.0 mL of unknown concentration of HCl into a 250 mL Erlenmeyer flask, add three drops of phenolphthalein into the acid.
3. Record the initial volume value for the NaOH, then add the NaOH drop by drop into the acid until you see a light pink colour.
4. Record the final volume of the NaOH
5. Dispose of the contents of the flask into the sink (with lots of water)
6. Continue with the next trial

Data Table:

Molarity of NaOH = 0.50 M	Trial 1	Trial 2	Trial 3	Trial 4
Initial volume				
Final volume				
Volume of used				

Use the average volume of NaOH from the 4 trials to calculate the **concentration** of the HCl in the space below and hand it in tomorrow!