

## CHEMISTRY 12

### TABLE OF SPECIFICATIONS FOR THE PROVINCIAL EXAMINATION

CURRICULUM			COGNITIVE LEVEL %			TOTAL %									
ORGANIZERS	SUB-ORGANIZERS	MARKS SUB-TOTAL	Knowledge	Understanding and Application	Higher Mental Processes										
<b>Reaction Kinetics</b>	A1, 2 Introduction	12	↑	↑	↑	12									
	A3, 4, 5 Collision Theory														
	A6, 7, 8 Reaction Mechanisms and Catalysts														
<b>Dynamic Equilibrium</b>	B1, 2 Introduction	8				↑	↑	↑	16						
	B3, 4 Le Châtelier's Principle	8													
	B5, 6 The Equilibrium Constant	8													
<b>Solubility Equilibria</b>	C1, 2, 3 Concept of Solubility	8							↑	↑	↑	16			
	C4, 5 Solubility and Precipitation	8													
	C6, 7, 8 Quantitative Aspects	8													
<b>Nature of Acids and Bases</b>	D1, 2, 3 Properties and Definitions	9										↑	↑	↑	9
	D4, 5, 6 Strong and Weak Acids and Bases														
<b>Acids and Bases: Quantitative Problem Solving</b>	E1, 2 $K_w$ , pH, pOH	8	↑	↑	↑										8
	E3, 4 $K_a$ and $K_b$ Problem Solving														
<b>Applications of Acid-Base Reactions</b>	F1 Neutralizations of Acids and Bases	16				↑	↑	↑							16
	F2, 3 Indicators														
	F4, 5 Hydrolysis of Salts														
	F6, 7 Buffer Solutions														
	F8 Acid Rain														
<b>Oxidation-Reduction</b>	G1, 2 Introduction	11							↑	↑	↑				11
	G3, 4 Balancing Redox Equations														
<b>Applications of Redox Equations</b>	H1, 2 Electrochemical Cells	12										↑	↑	↑	12
	H3 Corrosion														
	H4, 5 Electrolytic Cells														
<b>TOTAL %</b>			11	78	11	100									

The values in this table are approximate and may fluctuate.

Examination configuration: 62% in multiple-choice format  
38% in written-response format

#### Acknowledgement

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