Name:_____
Blk:_____Date:_____

Chemistry 12
ACID BASE PART II Lesson # 18
PRACTICAL ASPECTS OF TITRATIONS

To carry out a titration you must have a solution of _____concentration. This is also referred to as a _____ or a _____ _____.

A _____ _____ is a substance that is used to determine the concentration of a standard solution. A primary standard is one that can be obtained _____
_____.

There are TWO WAYS to prepare a **STANDARD SOLUTION**:
1. To prepare a standard solution of a base eg. NaOH
a.

b.

2.To prepare a standard solution of an acid eg. HCl

SEATWORK: Do Exercises 121-123 pg 165 in HEBDEN
PLO's: P1 (PRIMARY STANDARDS AND STANDARDIZED SOLUTIONS)

TYPES OF NEUTRALIZATION REACTIONS

Recall from earlier in this unit that we investigated the Formula, Complete and Net Ionic equations for a STRONG ACID and STRONG BASE neutralization reaction:
Ex. When NaOH reacts with HCl

i.

ii.

iii.

When the $[H_3O^+] = [OH^-]$ the solution is NEUTRAL, but if one is in excess it is either basic or acidic (depending on which ion is in excess)

When a WEAK ACID is reacted with a STRONG BASE:
eg. HF + NaOH

i.

ii.

iii.

When a WEAK BASE is reacted with a a STRONG ACID:
Ex. NH_3 + HCl

i.

ii.

iii.

SEATWORK/HOMEWORK: Worksheet
PLO's: P4

Name:_____
Blk:_____Date:_____

Chemistry 12
ACID BASE PART II Lesson # 18
PRACTICAL ASPECTS OF TITRATIONS WORKSHEET

Write out the FORMULA, COMPLETE and NET IONIC Equations for the following neutralization reactions:

1. $Sr(OH)_2 + HBr$

2. $NaCN + HCl$

3. $C_6H_5COOH + LiOH$

4. $HCl + KOH$

5. $Na_2SO_3 + HNO_3$

6. $CH_3COOH + RbOH$